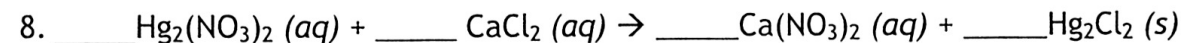
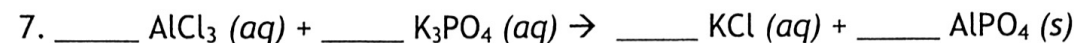
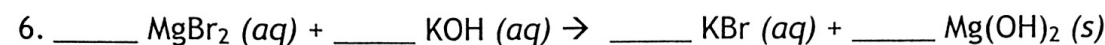
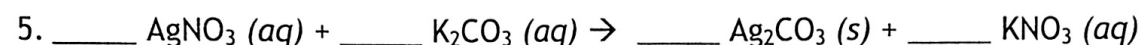
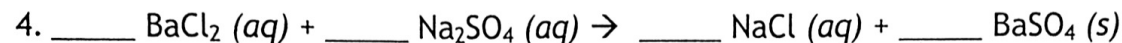
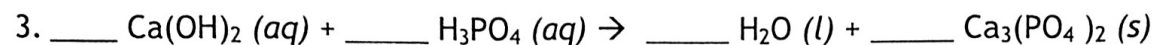
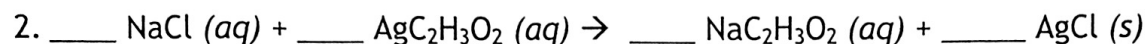
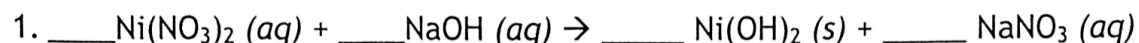


1. Name or give the chemical formula for each of the following compounds.
2. State whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution. Use solubility rules.

| Chemical Formula | Name | Solubility |
|---------------------------------------|---------------------|------------|
| 1. $\text{NH}_4\text{CH}_3\text{COO}$ | | |
| 2. $\text{Ba}(\text{OH})_2$ | | |
| 3. | Iron (II) Carbonate | |
| 4. NaOH | | |
| 5. RbNO_3 | | |
| 6. | Cesium Sulfate | |
| 7. MgSO_4 | | |
| 8. ZnCl_2 | | |
| 9. | Zinc Hydroxide | |
| 10. $\text{Zn}_3(\text{PO}_4)_2$ | | |
| 11. AgBr | | |
| 12. KNO_3 | | |
| 13. Al_2S_3 | | |
| 14. | Silver Acetate | |
| 15. Sr_2CrO_4 | | |
| 16. | Aluminum Phosphate | |
| 17. BaSO_4 | | |
| 18. $\text{Ca}(\text{OH})_2$ | | |
| 19. BaCO_3 | | |
| 20. MgCrO_4 | | |
| 21. | Iron (III) sulfide | |
| 22. NH_4CN | | |
| 23. | Silver Iodide | |
| 24. Hg_2SO_4 | | |

Net Ionic Equations Worksheet

BALANCE the following equations then write the NET IONIC EQUATION for each one:



Using the solubility rules, predict the products, balance the equation, and write the complete ionic and net ionic equations for each of the following reactions.

